

# Engineering Essentials Solution Guide

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Solution: (a)  $(1000 \times 10^3) (6.022 \times 10^{23} \text{ atoms/mol}) (55.847 \text{ g/mol})$  (b)  $(1 \text{ mol}) (10.81 \text{ g/mol}) = 4.6 \text{ cm}^3 \times 2.36 \text{ g/cm}^3 = 1.078 \times 10^{28} \text{ atoms}$

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